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$\text{CO} + \frac{1}{2} \text{O}_2 \rightarrow \text{CO}_2$ is a full-cell reaction common in SOFCs; the half-cell reactions would be $\text{O}_2 + \text{CO} \rightarrow \text{CO}_2 + 2 \text{e}^-$, which is an Oxidation, or anode, reaction, and $\frac{1}{2} \text{O}_2 + 2 \text{e}^- \rightarrow \text{O}^{2-}$ -which is a reduction, or cathode reaction. 2. $\frac{1}{2} \text{O}_2 + \text{H}_2 \rightarrow \text{H}_2\text{O}$ is a full cell reaction in SOFCs or PEMs, depending on the circulating ion (O^{2-} -or H^+ , respectively).

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Problem 1.8 (20 points) You can see the benefits in a fuel cell of the independent scaling of the fuel reservoir and the fuel cell stack. Therefore, for this problem, you may compute the volume of the stack and the reservoir independently. For the stack, you need to supply 30 kW with a fuel cell that supplies power at 1 kW/L and 500 W/kg.

Fuel Cell Fundamentals Solutions

The book also introduces the five types of fuel cells (PEMFC, PAFC, AFC, MCFC, SOFC) and gives the reader a good understanding of the characteristics, advantages and disadvantages to each one, allowing for easy comparison.

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